UNIT– I: INORGANIC CHEMISTRY

1. Atomic structure and periodic properties[10 Hours]

Dual nature of electron, De-Broglie’s equation, Heisenberg’s uncertainty principle, quantum numbers, Aufbauprinciple, Pauli’s exclusion and Hund’s rule for electron configuration.

Periodicity in atomic properties and its causes, explanation of general trends of periodic properties: atomic and ionic radii, ionization potential, electronegativity and electron affinity.

Chemistry of s and p block elements

Special characteristics such as metallic character, polarizing power, hydration energy, inert pair effect, relative stability of different oxidation state, diagonal relationship of (1) lithium with magnesium (2) boron with silicon and (3) beryllium with aluminum, anomalous behavior of Li and Be, formation of complex compounds, catenation, allotropy (diamond and graphite-theirstructure, properties and its uses).

2. Chemical bonding in covalent compounds [10 Hours]

Covalent bond: Valence bond theory and its limitations. Concept of hybridization: sp(BeCl₂), sp² (BF₃), sp³ (SiH₄), sp³d (PCl₅) and sp³d² (SF₆).

Stereochemistry of inorganic molecules: Sidgwick Powell rule and VSEPR theory. Structure of molecules: SnCl₂, SO₄²⁻, CO₃²⁻.

UNIT-II: ORGANIC CHEMISTRY

3. Basic organic chemistry and aliphatic hydrocarbons containing σ bond

[10 Hours]

Nomenclature of organic compounds (Only acyclic - IUPAC-1993).
Electronic displacements: Inductive, electromeric, mesomeric effects and hyper conjugation. Applications of inductive effect to bond length, dipole-moment, reactivity of alkyl halides, relative strength of acid, basicity of amines.

Homolytic and heterolytic fission, curly arrow rules
Reaction intermediates: carbocation, carbanion, free radical, carbenes and benzynes (formation by cleavage type, structure, relative stabilities, generation). Types of organic reagents: Nucleophiles and electrophiles.

Types of organic reactions: Substitution, addition, elimination and rearrangement. Nucleophilic substitution reaction mechanism (S_N1 & S_N2) for alkyl halides.


4. Aliphatic hydrocarbons (acyclic)[10 Hours]

Chemistry of alkanes:

Hydrocarbons containing carbon-carbon π bonds
Formation of alkene elimination reactions, dehydration of alcohols, dehydrohalogenation of alkyl halides, dehalogenation of vicinal and germinal dihalides.
Mechanism of E1, E2, E1cb reactions, Saytzeff and Hofmann eliminations. Electrophilic addition reactions and their mechanisms (Markownikov/ Anti Markownikov rule).

Reactions of alkenes: Oxymercuration-demercuration, hydroboration oxidation, ozonolysis, reduction (catalytic), syn and anti-hydroxylation (oxidation), 1,2-and 1,4-addition reactions in conjugated dienes, Diels-Alder reaction.
Formation of alkynes: Dehydrohalogenation of vicinal and geminal dihalides, dehalogenation of tetrahalides

Reactions of alkynes: Acidity, electrophilic addition reactions like halogenation, hydrohalogenation, hydration, hydroboration, addition of carbene and catalytic hydrogenation. Nucleophilic addition with hydrogen cyanide and alcohol, hydration to form carbonyl compounds, alkylation of terminal alkynes.

UNIT III: PHYSICAL CHEMISTRY

5. Chemical kinetics [12 Hours]

   Concept of chemical kinetic: rate of chemical reaction, concentration dependence of reaction rate specific reaction rate constant, order and molecularity of the reaction. Factors affecting rate of the reaction.

   Definition, derivation of integrated rate equations for zero, first and second (same and different reactants) order reactions, their characteristics and half-life periods.

   Determination of the order of reaction: (1) Hit and trial method or Integration method and its limitations (2) Fractional change method, (3) Oswald’s isolation method (4) graphical method and (5) van’t Hoff differential method. Concept of activation energy. Derivation of Arrhenius equation and determination of activation energy by integrated equation and graphical methods.

   Theories of reaction rates: Collision theory and absolute reaction rate theory of bimolecular reactions and qualitative comparison. Numericals

6. Adsorption [4 Hours]

   Introduction, types of adsorption (physical and chemical), characteristics and factors affecting adsorption. Adsorption isotherm and Freundlich equation. Langmuir theory of adsorption: assumptions, derivation, modification in equation at very low and high pressure, limitations and applications of adsorption.

7. Catalysis [4 Hours]

   Introduction, types of catalysis (homogeneous and heterogeneous), characteristics of catalysis, autocatalysis, negative catalysis (Inhibitor), promoters, and catalytic poisoning. Activation energy and catalysis. Theories of catalysis: (1) Intermediate
compound formation and (2) adsorption theory, active centers. Enzyme catalysis and its characteristics.

**Reference books: Inorganic Chemistry**

2. Inorganic Chemistry - J. N. Gurtu & H. C. Khera
3. Concise of Inorganic Chemistry - J. D. Lee.
5. Advanced Inorganic Chemistry - Raymond Chang
8. Atomic Structure and Chemical Bonding by Manas Chanda.

**Reference books: Organic Chemistry**

1. Organic Reaction Mechanism, including Reaction Intermediates , V. K. Ahluwalia, Ane's Chemistry active series.
10. Stereochemistry of Organic–Principles and Application by D. Nasipuri 3rd addition
11. Stereochemistry Conformation and Mechanism by P. S. Kalsi

**Reference books: Physical Chemistry**

SAURSHTRA UNIVERSITY
B. Sc. SEMESTER-I
C-102: CHEMISTRY PRACTICALS

3 Credits
50 Marks

1. Organic qualitative analysis [20 marks]
(Minimum 12 compounds)
Compounds containing one functional group such as phenolic, carboxylic acid, ester, amide, nitro, amine, aldehyde, ketone, alcohol, halogen, anilide, carbohydrate and hydrocarbon.

List of compounds: Benzoic acid, cinnamic acid, phenol, α-naphthol, β-naphthol, acetone, ethyl methyl ketone, methyl acetate, ethyl acetate, naphthalene, anthracene, aniline, nitrobenzene, benzamide, urea, thiourea, chloroform, acetonilide, carbon tetra chloride, chloro benzene, bromo benzene.

2. Volumetric analysis [15 Marks]

Part-1 Acid-base titrations

1. To prepare a solution by dissolving ‘x’ g NaHCO₃ /Na₂CO₃ in 100 ml solution and determine its concentration in terms of normality and molarity using 0.1 N HCl solution.
2. To determine the normality, molarity and g/lit of NaOH and HCl using 0.1 N Na₂CO₃ solution.
3. To determine the normality, molarity and g/lit of each component in a given mixture of NaHCO₃ and Na₂CO₃ using 0.1N HCl solution.

Part-2 Redox titrations

4. To determine the normality, molarity and g/lit of each component in a mixture of H₂C₂O₄.2H₂O and H₂SO₄ using 0.1 N KMnO₄ and 0.1 N NaOH solution.
5. To determine the normality, molarity and g/lit of each component in a mixture of H₂C₂O₄.2H₂O and K₂C₂O₄.H₂O using 0.1N NaOH and 0.1 N KMnO₄ solution
6. To determine the normality, molarity and g/lit of KMnO₄ and FeSO₄.7H₂O solution using 0.1 N H₂C₂O₄.2H₂O solution.
To determine the normality, molarity and g/lit of FeSO$_4$ (NH$_4$)$_2$SO$_4.6$H$_2$O and K$_2$Cr$_2$O$_7$ solutions using 0.1 N KMnO$_4$ solution.
UNIT-I: INORGANIC CHEMISTRY

1. Basics of ionic compounds [7Hours]
   Introduction, characteristics of ionic solids, Born Haber cycle and its applications, Max Born equation, limiting radius ratio. Relation between radius ratio, co-ordination number and crystal structure. Derivation of $r+/r-$ ratio in triangular, planar, square planar, body centered and tetrahedral crystal lattices. Defects in ionic crystals (stoichiometric and non-stoichiometric), study of N & P types of semi-conductors.

2. Chemistry of elements of 3d series [6 Hours]
   Introduction, definition, electronic configuration, reversal of energies of 3d and 4s orbitals, physical properties such as atomic properties (atomic radii, ionic radii, and ionization potential), metallic conductivity, melting point & boiling point, density, reducing properties, tendency of formation of alloys, catalytic properties, magnetic and spectral properties. Calculation of spin only magnetic momentum of inner orbital and outer orbital complexes $[\text{NiCl}_4]^2$, $[\text{Ni(CN)}_4]^2$, $[\text{FeF}_6]^{-4}$, $[\text{Fe(CN)}_6]^{4}$.

3. Basics of co-ordination chemistry [7Hours]
   Warner theory, types of ligands (simple ligands, π-acid ligands, according to number of donating electrons, chelating ligands) with definition and examples. Co-ordination number and geometry related to co-ordination number. Isomerism and its classification (structural and stereo isomerisms). In structural isomerism: (1) ionization(2) hydration (3) co-ordination (4) co-ordination positions (5) polymerization and (6) linkage isomerism. Geometric/cis-trans isomerism in $\text{ML}_4$ and $\text{ML}_6$ types of complexes.
UNIT II: ORGANIC CHEMISTRY

4. Cycloalkanes [10 Hours]

Introduction and classification of ring system (monocyclic and polycyclic, size, number of carbon atoms common between the two rings).

IUPAC nomenclature of cycloalkanes (including simple spiro compounds, fused ring and bridged ring systems—bicyclic only).

Method of preparation of small ring cycloalkanes: Intra-molecular Wurtz’s reaction, Simmons-Smith, Diels-Alder reaction.

Chemical properties of cycloalkanes: Substitution reactions, addition reactions, Baeyer’s strain theory and its limitations (puckering).

Conformations and conformational analysis: Conformation of ethane, propane and butane.

5. Aromatic hydrocarbons [10 hours]

Aromaticity: Criteria for aromatic, non-aromatic and antiaromatic types, applications of Hückel’s rule to simple annulene, cyclic carbocation/anion.

Electrophilic aromatic substitution reactions of benzene with mechanism. Theory of effect of substituents on reactivity and orientation (with resonating structures for activating and deactivating groups).

Electrophilic aromatic substitution reactions of the followings with mechanisms: Halogenation, nitration, sulphonation, Friedel Crafts alkylation, Friedel Crafts acylation.

UNIT III: PHYSICAL CHEMISTRY

6. Ionic equilibria [12 Hours]

Types of electrolytes. Degree of ionization and factors affecting degree of ionization. Ionization constant and Ionic product of water. Ionization constants of weak acids and bases. Common ion effect and calculation of concentrations of OH⁻ ions (NH₄Cl+NH₄OH) and H⁺ ions (H₂S+HCl). Solubility and solubility products of sparingly soluble salts—applications of solubility product.
Hydrolysis of salts: Definition of hydrolysis of salts. Salts of strong acids and bases. Relation among $K_h$, $K_a$, or $K_b$ and $K_w$. Degree of hydrolysis and pH of the solution of salts of weak acids and strong bases, salts of weak bases and strong acids and salts of weak bases and weak acids.


7. **Solid state** [8 Hours]


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11. Stereochemistry Conformation and Mechanism by P. S. Kalsi

**Reference books: Physical Chemistry**

1. **Qualitative analysis of inorganic salts:** [20 Marks]
   (Minimum 12 salts-containing two radicals)

   Inorganic salts containing anion(chloride, bromide iodide, nitrate, nitrite, sulphate, sulphite, sulphide, carbonate, phosphate (soluble & insoluble), oxide, chromate, and dichromate).

2. **Inorganic volumetric analysis** [15 Marks]
   (Standard solution should be given)

   1. Quantitative estimation of Cu\(^{2+}\) in a given CuCl\(_2\).2H\(_2\)O solution using 0.01M EDTA solution.
   2. Quantitative estimation of Ni\(^{2+}\) in a given NiSO\(_4\).7H\(_2\)O solution using 0.01M EDTA solution.
   3. Quantitative estimation of Zn\(^{2+}\) in a given ZnCl\(_2\) solution using 0.01M EDTA solution.
   4. Determination of total hardness of water by EDTA.
   5. Determination of acetic acid in a commercial vinegar using 0.1M NaOH solution.
   6. Determination of alkali in antacid using 0.1M HCl solution.
   7. Quantitative estimation of Fe\(^{2+}\) by dichromate method (Internal indicator method).
INSTRUCTIONS

1. B. Sc. Chemistry Syllabus for Semester I & II consists of three units:
   UNIT – I: Inorganic Chemistry
   UNIT – II: Organic Chemistry
   UNIT – III: Physical Chemistry
2. All units carry equal weightage.
3. All units should be given equal weightage in the question paper.
4. 70 Marks for Semester & 30 marks for Internal Examinations.
5. Time duration: 2 1/2 Hours
6. Question 1 carries 20 Marks
7. Questions 2 & 3 carry 25 Marks each

Question 1: Shorts questions (From UNIT I-III) [20 Marks]
   (One word, one line, explain, definition, true or false, fill up the blanks, MCQ, etc.)

Question 2: From Units I-III [25 marks]
   a. Answer Any 3 out of 4 (2 questions from Unit I and one question from Unit II & III). [2x3=6 Marks]
   b. Answer Any 3 out of 4 (2 questions from Unit II and one question from Unit I & III). [3x3=9 Marks]
   c. Answer Any 2 out of 3 (1 question from Unit I to III). [5x2=10 Marks]

Question 3: From Units I-III [25 marks]
   a. Answer Any 3 out of 4 (2 questions from Unit III and one question from Unit I & II). [2x3=6 Marks]
   b. Answer Any 3 out of 4 (2 questions from Unit II and one question from Unit I & III). [3x3=9 Marks]
   c. Answer Any 2 out of 3 (1 question from Unit I to III). [5x2=10 Marks]